

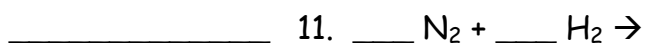
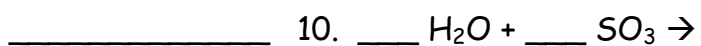
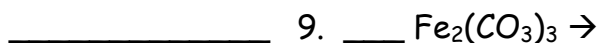
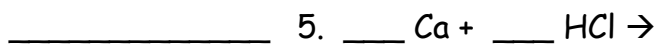
Reactions Worksheet

Name _____ Period _____ Date _____

I] Name, predict, and balance each of the following: (included are synthesis (S), decomposition (D), combustion (C), single replacement(SR), acid base(AB), and double replacement(DR).

NAME of RXN	PREDICT / BALANCE
_____	1. ____ CH_3COCH_3 + ____ $\text{O}_2 \rightarrow$
_____	2. ____ C_2H_2 + ____ $\text{O}_2 \rightarrow$
_____	3. ____ Na + ____ $\text{O}_2 \rightarrow$
_____	4. ____ Al_2O_3 + ____ $\text{CO}_2 \rightarrow$
_____	5. ____ CaO + ____ $\text{H}_2\text{O} \rightarrow$
_____	6. ____ SO_3 + ____ $\text{H}_2\text{O} \rightarrow$
_____	7. ____ $\text{HgO} \rightarrow$
_____	8. ____ $\text{CuCO}_3 \rightarrow$
_____	9. ____ $\text{Fe}(\text{OH})_3 \rightarrow$
_____	10. ____ $\text{H}_2\text{CO}_3 \rightarrow$
_____	11. ____ Zn + ____ $\text{CuSO}_4 \rightarrow$
_____	12. ____ Al + ____ $\text{SnCl}_2 \rightarrow$
_____	13. ____ NaBr + ____ $\text{Cl}_2 \rightarrow$
_____	14. ____ H_2SO_4 + ____ $\text{NH}_4\text{OH} \rightarrow$
_____	15. ____ FeBr_2 + ____ $\text{K}_2\text{CO}_3 \rightarrow$
_____	16. ____ $\text{Ba}(\text{NO}_3)_2$ + ____ $\text{Na}_2\text{SO}_4 \rightarrow$
_____	17. ____ LiOH + ____ $\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow$

II.] Additional Practice: name, predict, and balance.



Reactions Worksheet

Name _____ Period _____ Date _____

I.] Name, predict, and balance each of the following: (included are synthesis (S), decomposition (D), combustion (C), single replacement(SR), acid base(AB), and double replacement(DR).

NAME of RXN	PREDICT / BALANCE
<u>Combustion</u>	1. $\text{CH}_3\text{COCH}_3 + \text{\textcolor{red}{4}} \text{O}_2 \rightarrow \text{\textcolor{red}{3}} \text{CO}_2 + \text{\textcolor{red}{3}} \text{H}_2\text{O}$
<u>Combustion</u>	2. $\text{\textcolor{red}{2}} \text{C}_2\text{H}_2 + \text{\textcolor{red}{5}} \text{O}_2 \rightarrow \text{\textcolor{red}{4}} \text{CO}_2 + \text{\textcolor{red}{2}} \text{H}_2\text{O}$
<u>Synthesis</u>	3. $\text{\textcolor{red}{2}} \text{Na} + \text{\textcolor{red}{2}} \text{O}_2 \rightarrow \text{\textcolor{red}{2}} \text{Na}_2\text{O}$
<u>Synthesis</u>	4. $\text{\textcolor{red}{2}} \text{Al}_2\text{O}_3 + \text{\textcolor{red}{3}} \text{CO}_2 \rightarrow \text{\textcolor{red}{2}} \text{Al}_2(\text{CO}_3)_3$
<u>Synthesis</u>	5. $\text{\textcolor{red}{2}} \text{CaO} + \text{\textcolor{red}{2}} \text{H}_2\text{O} \rightarrow \text{\textcolor{red}{2}} \text{Ca}(\text{OH})_2$
<u>Synthesis</u>	6. $\text{\textcolor{red}{2}} \text{SO}_3 + \text{\textcolor{red}{2}} \text{H}_2\text{O} \rightarrow \text{\textcolor{red}{2}} \text{H}_2\text{SO}_4$
<u>Decomposition</u>	7. $\text{\textcolor{red}{2}} \text{HgO} \rightarrow \text{\textcolor{red}{2}} \text{Hg} + \text{\textcolor{red}{O}_2}$
<u>Decomposition</u>	8. $\text{\textcolor{red}{2}} \text{CuCO}_3 \rightarrow \text{\textcolor{red}{2}} \text{CuO} + \text{\textcolor{red}{2}} \text{CO}_2$
<u>Decomposition</u>	9. $\text{\textcolor{red}{2}} \text{Fe}(\text{OH})_3 \rightarrow \text{\textcolor{red}{Fe}_2\text{O}_3} + \text{\textcolor{red}{3}} \text{H}_2\text{O}$
<u>Decomposition</u>	10. $\text{\textcolor{red}{2}} \text{H}_2\text{CO}_3 \rightarrow \text{\textcolor{red}{2}} \text{H}_2\text{O} + \text{\textcolor{red}{2}} \text{CO}_2$
<u>S. Replacement</u>	11. $\text{\textcolor{red}{Zn}} + \text{\textcolor{red}{CuSO}_4} \rightarrow \text{\textcolor{red}{ZnSO}_4} + \text{\textcolor{red}{Cu}}$
<u>S. Replacement</u>	12. $\text{\textcolor{red}{2}} \text{Al} + \text{\textcolor{red}{3}} \text{SnCl}_2 \rightarrow \text{\textcolor{red}{2}} \text{AlCl}_3 + \text{\textcolor{red}{3}} \text{Sn}$
<u>S. Replacement</u>	13. $\text{\textcolor{red}{2}} \text{NaBr} + \text{\textcolor{red}{Cl}_2} \rightarrow \text{\textcolor{red}{2}} \text{NaCl} + \text{\textcolor{red}{Br}_2}$
<u>D.R. / A-B</u>	14. $\text{\textcolor{red}{H}_2\text{SO}_4} + \text{\textcolor{red}{2}} \text{NH}_4\text{OH} \rightarrow \text{\textcolor{red}{(NH}_4)_2\text{SO}_4} + \text{\textcolor{red}{2}} \text{H}_2\text{O}$
<u>D. Replacement</u>	15. $\text{\textcolor{red}{FeBr}_2} + \text{\textcolor{red}{K}_2\text{CO}_3} \rightarrow \text{\textcolor{red}{FeCO}_3} + \text{\textcolor{red}{2}} \text{KBr}$
<u>D. Replacement</u>	16. $\text{\textcolor{red}{Ba}(\text{NO}_3)_2} + \text{\textcolor{red}{Na}_2\text{SO}_4} \rightarrow \text{\textcolor{red}{BaSO}_4} + \text{\textcolor{red}{2}} \text{NaNO}_3$
<u>D. Replacement</u>	17. $\text{\textcolor{red}{2}} \text{LiOH} + \text{\textcolor{red}{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2} \rightarrow \text{\textcolor{red}{2}} \text{LiC}_2\text{H}_3\text{O}_2 + \text{\textcolor{red}{Fe}(\text{OH})}_2$

III.] Additional Practice: name, predict, and balance.

